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Study on the inclusion Interactions of β-Cyclodextrin with Rhodamine B Base

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Abstract: Addition of β-Cyclodextrin to aqueous solutions of rhodamine B base results in the deaggregation of the dye to its monomer form. This is due to the association of monomeric rhodamine B base to the cyclodextrin. The absorption and emission spectra of the aqueous dye solution have been described. This has also been proved by UV-Vis, FTIR and Scanning Eelectron Microscope (SEM). Ground state (Kg) and excited state (Ke) formation constants have been calculated and reported.
Keywords: Rhodamine B base, β-Cyclodextrin, Fluorescence, FTIR.

1. INTRODUCTION:

Rhodamine dyes are extensively used as laser dyes. They have often been used to characterize novel laser system, examine the efficiency of various pump sources and obtain high-power lasers [1]. These dyes are commonly used in alcoholic solutions, even though the thermal properties of water are superior to those of any alcohol. Specifically, the variation of the refractive index of water with temperature is much smaller than that of ethanol [2]. This characteristic is particularly important for the development of high–power laser and for continuous-wave laser. The main reason for not using aqueous dye solutions is the extensive aggregation of the dye leading to the formation of dimers or higher aggregates [1,3,4]. Such aggregates quench internally the dye fluorescence and prevent effective lasing. This difficultly has been dealt with in the past by adding detergents or using solvent mixtures, leading in fact to laser action from such modified solutions. Yet, the concentrations of the additives are often in the range of 4-24%, high enough to adversely affect the superior thermal properties of the aqueous media [2].

Cyclodextrins (CD) are cyclic polysugars composed of glucose units linked by 1-4 glycoside bonds [5]. The hydrophobic cavity present in the CD structure is capable of binding organic substrates including dye molecules [6,7]. The fluorescence properties of dye are affected when associated with CD, and it has been shown that inclusion complexes exhibit considerable increase in the fluorescence quantum yields [8-10]. The decreasing of fluorescence intensity of the butyl rhodamine B (BRB) was studied by Xiashizhu et al. [11]. Either monomer or dimer emission can be enchanced depending on the size of the cavity. Recently, it has been demonstrated that dimer formation of thionine in aqueous solution is prevented in the presence of cyclodextrins [12]. This has been attributed to the association of thionine

monomer to the CD cavity maintaining a low concentration of the free dye form in water, at which the monomer predominates.

2. EXPERIMENTAL

2.1 Materials

All the reagents used were analytical ones. Prior to their use the purity of the organic solvents (Water, DMF & DMSO) was checked via fluorescence. βCyclodextrin and Rhodamine B base were purchased from Sigma-Aldrich, Bangalore. Double – distilled water was used throughout. Stock aqueous solutions of the βCyclodextrin (0.002M – 0.012M) were prepared daily and maintained at room temperature for use.

Stock standard solution of Rhodamine B base was prepared in 5 x 10^{-5}M concentration. Working solutions were analyzed before each measurement of fluorescence, and the UV-VIS spectrum from 200 to 800nm was recorded.

2.2 Apparatus

Fluorescence spectra were recorded using a JASCO model FP~550 spectrofluorometer, the light source of which is a 150W xenon lamp. Absorption spectra were obtained using a JASCO – UVIDEC – 650 Spectrophotometer. The IR spectra were recorded on an AVATAR – 360 Series FTIR spectrometer. The Rh B base fluorescence intensity was measured at the maximum emission wave length of 590nm, 563nm and 570nm after excitation of solutions at 553nm,531nm, 522nm in water, DMF, and in DMSO respectively. The study of Microscopic morphological structure and measurements were made with the JEOL JSM 5610 LV scanning election microscope (SEM).

2.3 Preparation of Sample for SEM

Accurately weighed 0.012M βCD was added to1ml of double distilled water and stirred over an electromagnetic stirrer until it was dissolved. The βCD solution was then slowly poured into 5 x 10^{-5} M concentrated Rhodamine B base solution. The above mixed solutions were continuously stirred for 48 hours at room temperature. The reaction mixture was put to dry in an oven at 150°C for 12 hours to obtain the powder product, which is the inclusion complex of Rhodamine B base and βCD used in the study.

3. RESULTS AND DISCUSSION

3.1 Effect of Solvents

The absorption and fluorescence spectra of Rhodamine B base were studied in different solvents and the experimental results were compiled and were presented **in Table 1.** The absorption spectra of Rhodamine B base in all solvents consist of absorption bands of longer wavelengths. H-O transition may account for the display by the compound of a broad visible band of considerable charge transfer (CT) on emission.

The absorption spectrum of Rhodamine B base in water is largely red - shifted than in other solvents DMF and DMSO but their spectra do not exhibit marked changes in their shape.

Fig.1. depicts the fluorescence spectra of Rhodamine B base in water and solvents DMF and DMSO and the compiled relevant data have been presented in **Table 1**.The large red shift in fluorescence when the width of the band is maximum and of increased polarity.

[DMF (6.4) < DMSO (7.2) < water (9)]

It is evidence of the hydrogen bonding tendency of the solvents.

With different solvents a large Stokes- shift emission band was observed and the sensitivity of the band to change in solvent polarity leads to the inference that greater charge transfer takes place.

3.2 Absorption and Fluorescence Spectral Studies

Fig.2 show the absorption spectra of Rhodamine B base in the absence and presence of different concentrations of cyclodextrin with absorption maxima at 553nm in water[535.50nm in DMF and 540nm in DMSO respectively]. A red shift observed in the presence of βCD leads the inference that βCD is interacting with Rhodamine B base.

The effect of βCD on the fluorescence spectra of Rhodamine B base is more pronounced than its effect in absorption spectra. Only the maximum intensity of fluorescence decreases and no shift (Red or blue) was observed. There is decrease in the intensity of fluorescence with the addition of βCD up to a concentration of 12 x 10^{-3} mol dm^{-3}. Fluorescence intensity of Rh B base is higher in DMSO (664 A.U) than in solutions of Water (520A.U) and DMF (468A.U). Quenching of fluorescence intensity suggests the formation of the inclusion complex of Rhodamine B base and βCD. A decrease in the fluorescence intensity on the formation of an inclusion complex was observed earlier [12].

Data present in Table 1 depict the absorption and emission maxima of Rhodamine B base, in βCD solutions. Absorption peak was observed at 553nm.

It can be seen that fluorescence characteristics of Rhodamine B base in the solvents water, DMF and DMSO undergo drastic changes in the presence of βCD (**Fig.3a,3b and 3c).**

Table – 1 : Absorption, log ε, fluorescence Spectral data (nm) and Stoke's shift (cm^{-1}) of Rh B base in different solvents

Solvents	λ_{abs} (nm)	log ε ($M^{-1}cm^{-1}$)	λ_{flu} (nm)	stokes shift (cm^{-1})
Water	553	8.043	590	1134
DMF	531	8.026	565	1133
DMSO	522	8.018	570	1613

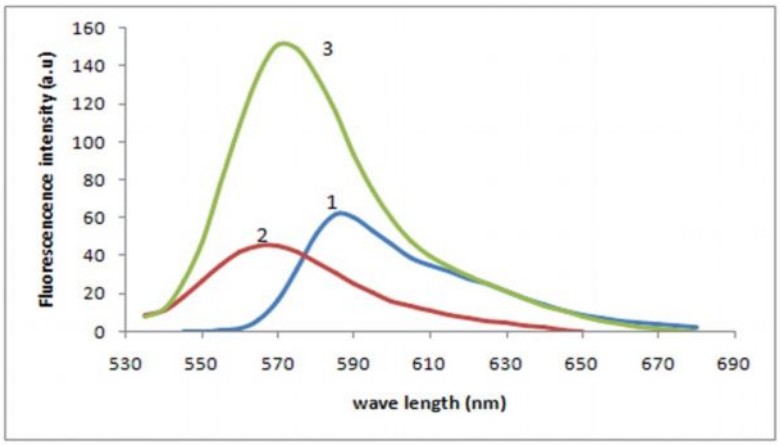

Fig. 1: Fluorescence spectra of Rh B base in different solvents
1. Water, 2.DMF, 3. DMSO

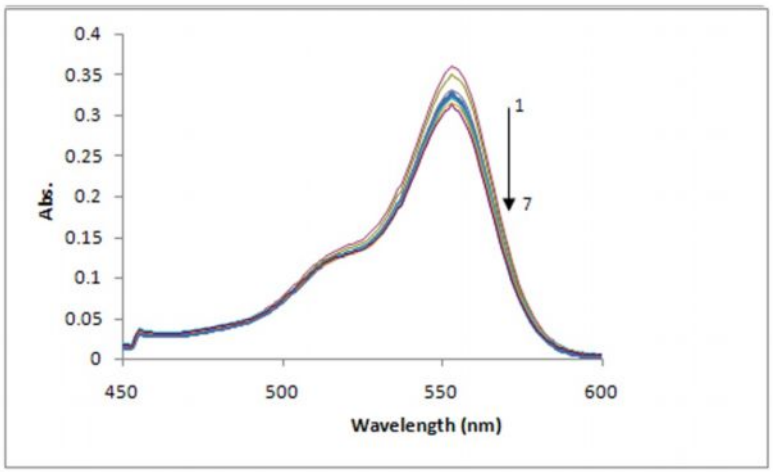

Fig. 2: Absorption spectra of Rh B base in different βCD concentrations (mol dm^{-3}) in water.
(1) 0, (2) 0.002 (3) 0.004, (4) 0.006, (5) 0.008, (6) 0.010, (7) 0.012

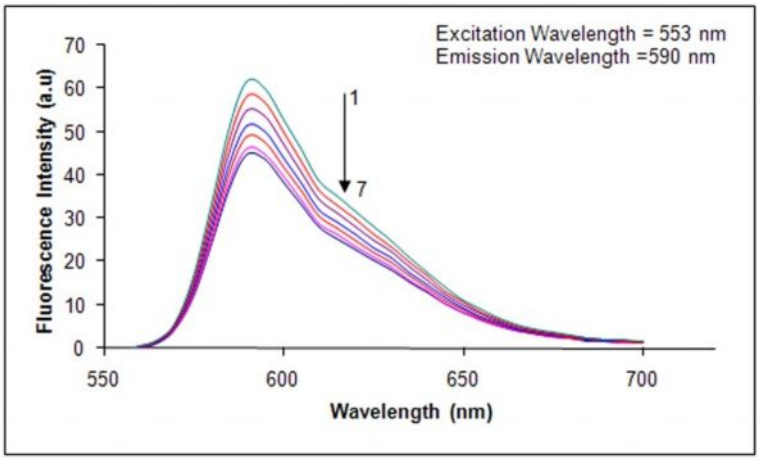

Fig. 3a: Fluorescence spectra of Rh B base in different βCD concentrations of (mol dm^{-3}) in Water
(1) 0, (2) 0.002 (3) 0.004, (4) 0.006, (5) 0.008, (6) 0.010, (7) 0.012

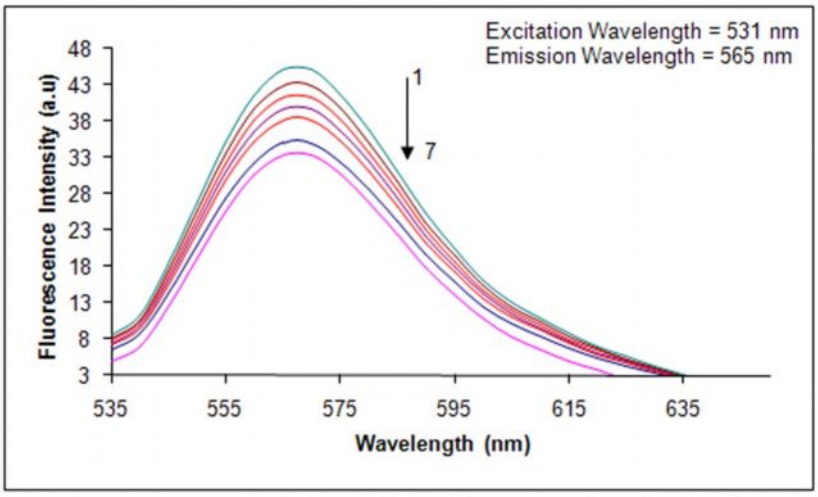

Fig. 3b: Fluorescence spectra of Rh B base in different βCD concentrations of (mol dm^{-3}) in DMF
(1) 0, (2) 0.002 (3) 0.004, (4) 0.006, (5) 0.008, (6) 0.010, (7) 0.012

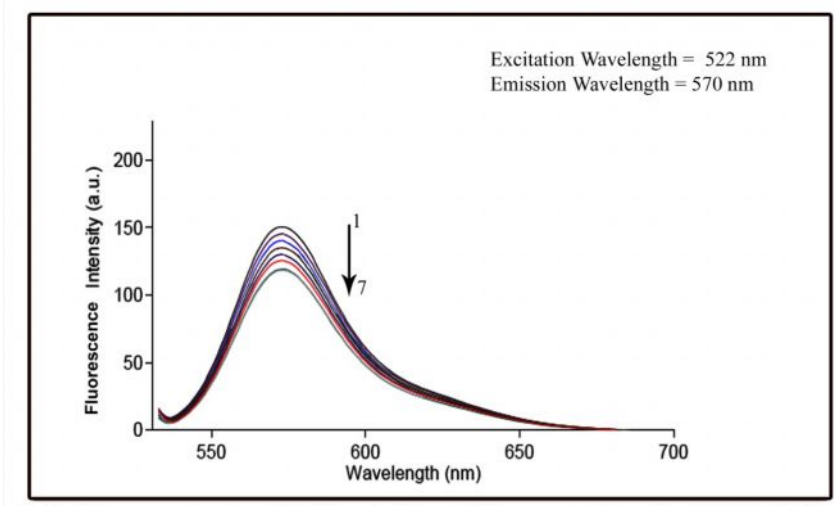

Fig. 3c: Fluorescence spectra of Rh B base in different βCD concentrations of (mol dm^{-3}) in DMSO
(1) 0, (2) 0.002 (3) 0.004, (4) 0.006, (5) 0.008, (6) 0.010, (7) 0.012

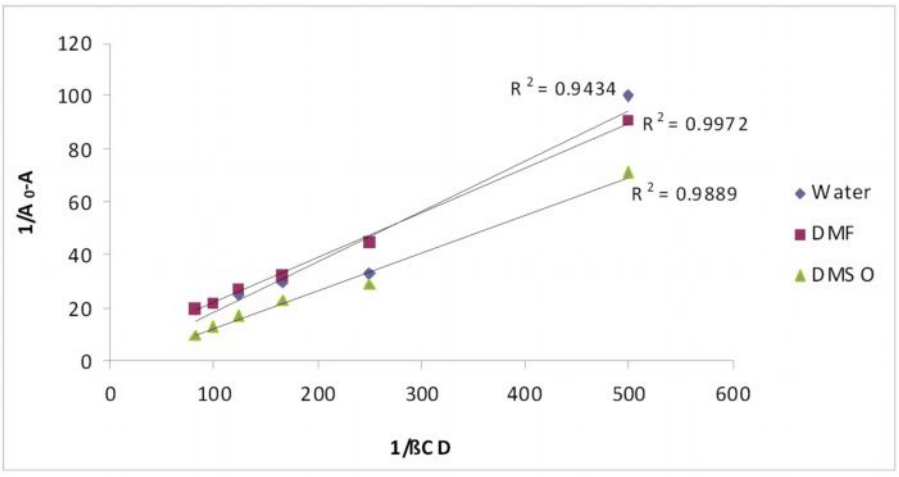

Fig. 4. : Plot of $(1/A_0 - A)$ and $(1/\beta\ CD)$ for Rh B base

Table – 2 : Formation constant k (M^{-1}) and free energy ΔG ($kJmol^{-1}$) of Rh B base with βCD

Solvent	βCD			
	Kg (M^{-1})	Ke (M^{-1})	ΔGg(KJ mol^{-1})	ΔGe(KJ mol^{-1})
Water	0.129	0.0004	5.159	19.709
DMF	0.16	0.0496	4.616	7.566
DMSO	0.245	0.0003	3.543	20.434

The formation constant of the βCD: Rhodamine B base complex was determined analyzing changes in the maximum of the intensity of absorption and fluorescence with different concentrations of βCD. Using Benesi – Hildebrand relation[13] the formation constant of the complex (K) was determined; it indicates the formation of 1:1 complex of Rhodamine B base and βCD; the equilibrium can be written as,

Rh B base + βCD $\rightleftharpoons$ Rh B base: βCD[14],

where Rh B base, βCD and Rhodamine B base: βCD represent, Rhodamine B base, beta cyclodextrin and the 1:1 inclusion complex of βCD and Rh B base respectively.

Plotting $\left(\frac{1}{Ao - A}\right)$ and $\left(\frac{1}{\beta - CD}\right)$ will result in a straight line as **in fig.4.**The ground state formation constant (Kg) has been evaluated from the slope values of this graph, and been tabulated and presented in **Table 2. The** formation constants in DMSO are considerably higher in value than similar constants in water and DMF solvents. This is due to the dipole moment values of the solvents. The dipole moment of water is 1.85D,DMF is 3.82D and DMSO is 3.96D.

With increasing concentration of βCD the resulting straight line plotted with $\left(\frac{1}{I0 - I}\right)$ and $\left(\frac{1}{\beta - CD}\right)$ changes with the intensity of fluorescence (Fig.5). The excited state formation constant (Ke) has been evaluated from the slope values.

Free energy change can be calculated from the formation constant 'K' with the equation

$\Delta G = - RT \ln K$

The change in free energy ΔG, (ΔGg -ground state, ΔGe-excited state) increases according to the increase of the dielectric constant of the solvents.The values have been presented in **Table 2.**

The formation constant of the absorption spectrum of the dimer of Rhodamine was evaluated and the geometric structure of the aggregate was determined using exciton theory[15]. Deaggregation in concentrated solutions and variable complex formation in dilute solutions were proposed to account for the β-cyclodextrin induced effects[16].

The large red shifted emission maxima in all solvents would indicate a dipolar interaction between the molecules of the solute and solvent. The nature of emission is not always easy to ascertain, since it can be the result of a variety of causes including dimer formation (or other kinds of aggregates) in the ground state excimer emission or charge transfer processes[14,17]. In the case of Rhodamine B base, there may be the possibility of the formation of dimers.

Change in the free energy of the complex is higher in its excited state than in its ground state in all the three solvents viz., water, DMF and DMSO, which indicates the more strong inclusion complex was formed in excited state than in ground state. It can be seen that the difference in change in free energy between its ground and excitation states of the complexes in the three different solvents is in the following order:

DMSO > Water > DMF.

Table – 3: Fluorescence lifetime and amplitude of Rh B base with βCD

Sample	Life time (Sec.) τ	Relative amplitude	χ^2	Standard deviation
Rh B base	1.47×10^{-9} 2.65×10^{-9}	92.42 7.58	1.09	1.87×10^{-4} 1.00×10^{-4}
Rh B base + βCD	8.528×10^{-10} 2.01×10^{-9}	79.04 20.96	1.11	2.18×10^{-4} 8.35×10^{-5}

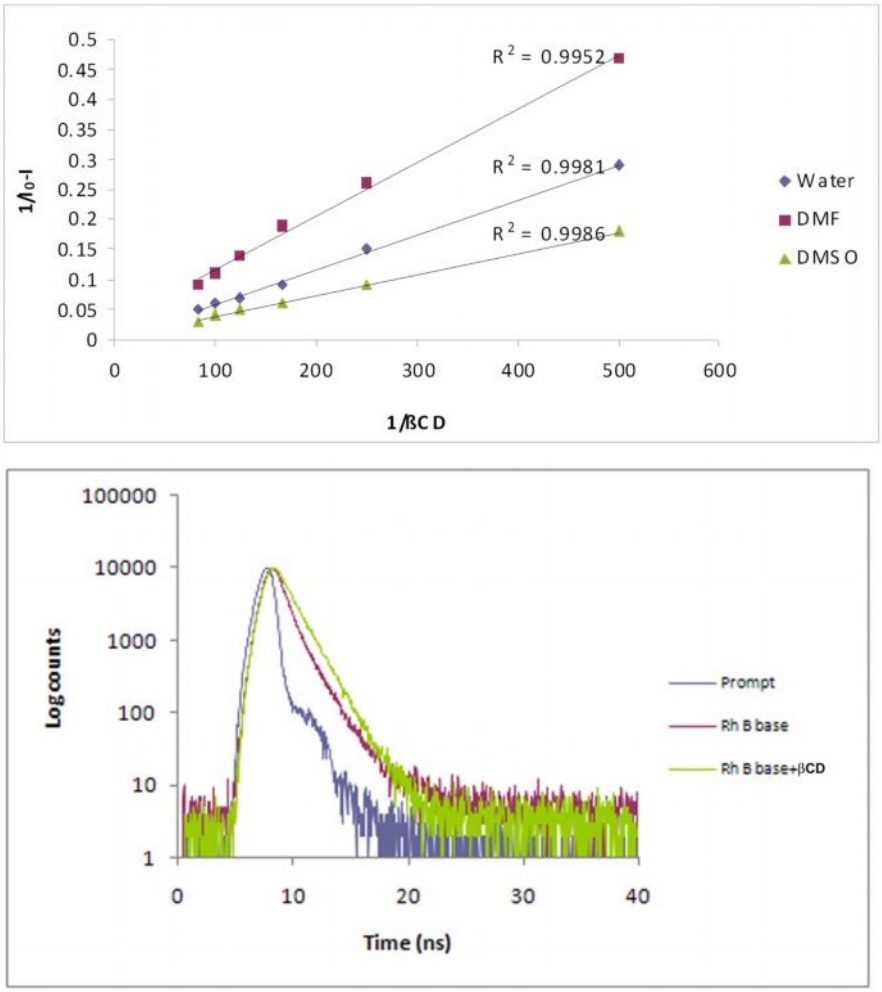

Fig. 6. Fluorescence decay curves of Rh B base and the inclusion complex of Rh B base with βCD

3.3 Life Time Spectral Study

The steady-state fluorescence and absorption measurement is not sufficient enough to have a clear picture of the formation of inclusion complex of Rh B base and βCD derivatives. Thus to have a better knowledge of the formation of inclusion complex, time resolved fluorescence experiments were performed for Rh B base and also for inclusion complexes. [Rh B base+ βCD]. **Fig. 6 gives** the fluorescence decay curves of (i) Rh B base and (ii) the inclusion complex of Rh B base and βCD. The Lifetimes and amplitudes of Rh B base with and without βCD derivatives are given **Table. 3.** The time – resolved fluorescence of Rh B base with βCD derivatives shows bi-exponential decay indicating the equilibrium between free and complexed forms. The χ^2 values for the single bi-exponential fitting are less than 1.5.

3.4 FTIR Spectral Study

The FTIR spectra of Rhodamine B base and the solid inclusion complex are shown in **Figs.7a and 7b respectively** and the values have been presented in **Table 4.** The comparison of Rhodamine B base with rhodamine 123 non-radiative rate constants indicated a relative decrease, suggesting that C-H and other lower frequency modes of vibrational energy transfer to solvent modes are less efficient than N-H stretching modes[18]. The absorption intensity of the inclusion complex is significantly weaker than that of the intensity of Rhodamine B base. The inclusion complex IR peaks in the range 2000 cm^{-1}--3000 cm^{-1} and is 1-3% weaker than that of the Rhodamine B base molecule. As there is no change in the wavelength other than the change in absorption intensities, it can be concluded that 1-3% weaker inclusion complex was formed of Rhodamine B base and βCD.

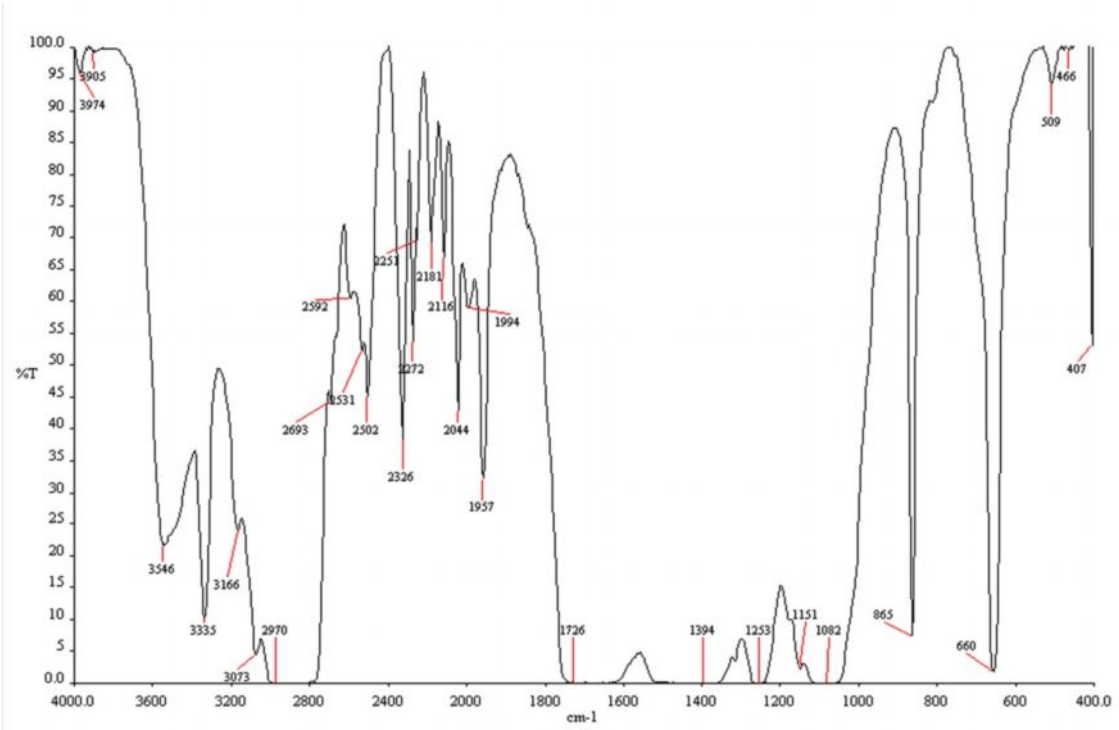

Fig. 7a : FTIR spectrum of Rh B base

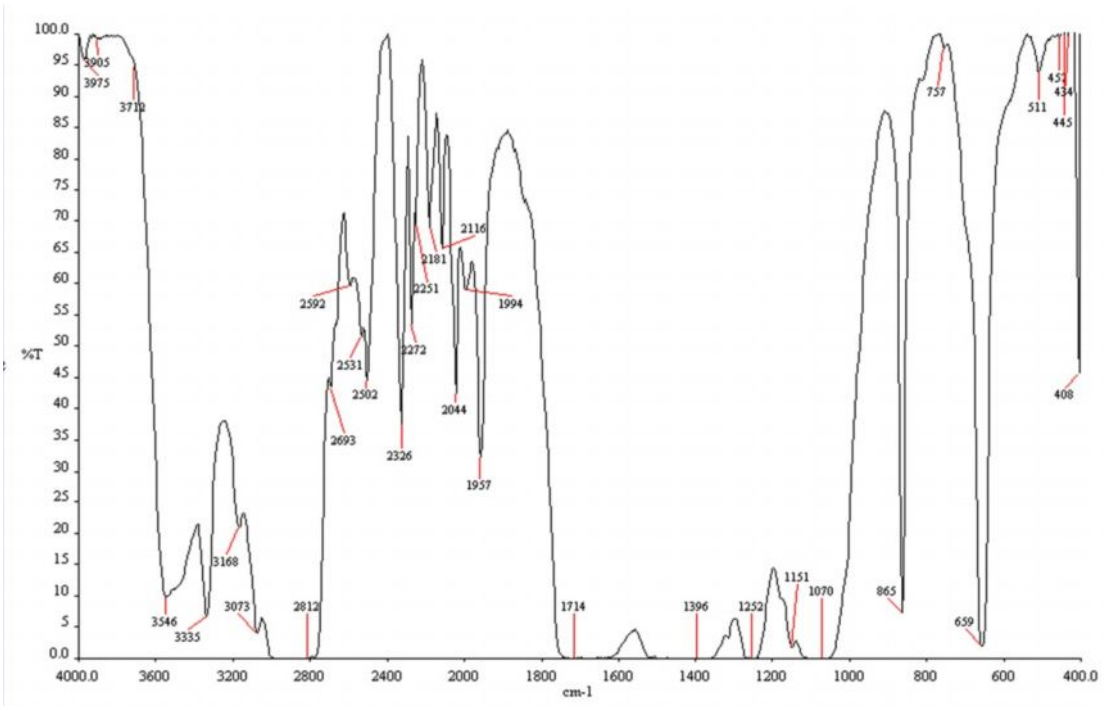

Fig. 7b : FTIR spectrum of (Rh B base + β CD) complex

Table –4 :Difference in FTIR absorption Peak intensities of Rh B base before and after the formation of inclusion complex

Tentative assignment	Rh B base (cm^{-1})	Inclusion complex (βCD) (cm^{-1})	Difference in intensitises prior to and after (%)
Hydroxy stretching	2693	2694	1
S – H stretching	2593	2593	3
S – H stretching	2531	2530	2
S – H stretching	2503	2502	3
C = N stretching	2326	2326	2
C = N stretching	2272	2272	1
C = N stretching	2251	2251	1
N = C = N antisym stretching	2181	2181	2

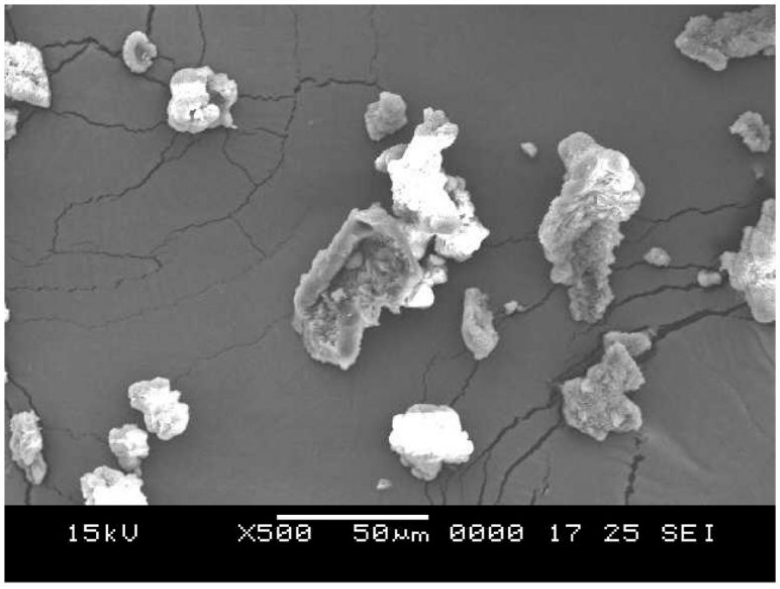

Fig. 8a Scanning electron microscope photograph of Rh B base

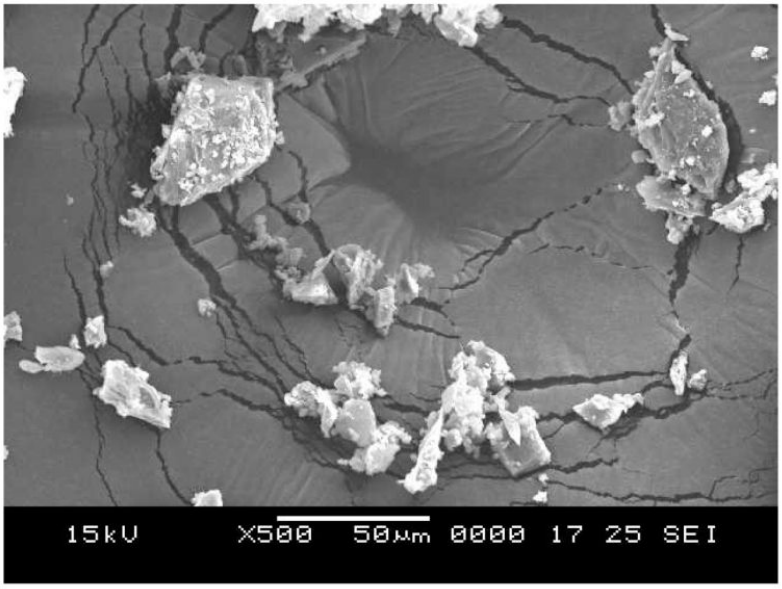

Fig. 8b Scanning electron microscope photograph of β CD

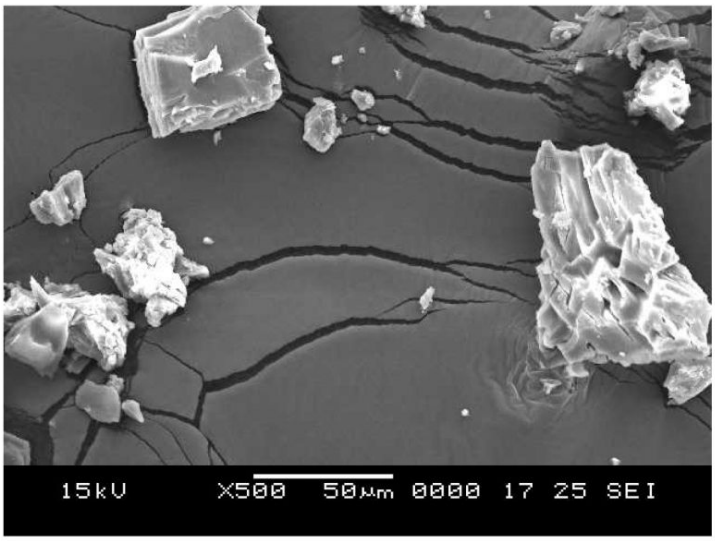

Fig. 8c Scanning electron microscope photograph of (Rh B base + β CD) complex

3.5 Scanning Electron Microscopic Studies

The powdered form of Rhodamine B base, βCD and the powdered form of the inclusion complex were observed through the scanning electron microscope and what was observed has been given in **figs.8a,8b and 8c .**It could be seen that the structure of the inclusion complex is different from that of the Rhodamine B base and βCD which is enough proof of the formation of a new inclusion complex.

4. CONCLUSION

The study demonstrates the effects of βCD on the photophysical properties of rhodamine B base in water , DMF and DMSO. Association of monomer Rhodamine B base to the hydrophopic cavity of β-Cyclodextrin induces the dissociation of dimers to the monomer dye forms. FTIR,and SEM results suggest that Rhodamine B base formed a solid inclusion complex with βCD.

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